1. **Use the Table of Ion Charges to name the following compounds:**

**eg. PbCO3  : Pb = Lead, CO3 = Carbonate ... PbCO3 is Lead Carbonate**

1. KNO3
2. FeSO4
3. NH4OH
4. Fe2(SO4)3
5. HgCl2
6. MgSO4
7. **How many atoms of oxygen are there in:**
8. 1 molecule of SO2
9. 3 molecules of H2SO4
10. 5 "molecules" of Fe3(PO4)2
11. 100 molecules of CH3COOH
12. 10 "molecules" of Ca(HCO3)2
13. **List the number and each kind of atom in each of:**
14. 3 molecules of ethanol, C2H5OH
15. 3 "molecules" of calcium phosphate, Ca3(PO4)2
16. 4 "molecules" of chromium (III) sulphate, Cr2(SO4)3
17. 2 "molecules" of magnesium chlorate, Mg(ClO3)2
18. **Use the table of Ion Charges on the previous page to write the chemical formulas for:**

**eg. Zinc Bromide:**

**Zinc Zn 2+, Bromide Br 1-**

**ZnBr2**

1. Potassium sulphate
2. Calcium carbonate
3. Magnesium oxide
4. Copper (I) chloride
5. Ammonium hydroxide
6. Barium nitrite
7. Silver chloride
8. Sodium nitrate
9. **Copy out the following table and fill in the blanks using your periodic table:**

|  |  |  |  |
| --- | --- | --- | --- |
| **Atomic Number** | **Element** | **Symbol** | **Possible Ion it forms** |
| **4** | **Beryllium** | **Be** | **Be2+** |
| **13** |  |  |  |
| **35** |  |  |  |
| **37** |  |  |  |
|  |  | **Ba** |  |
|  |  |  | **At -** |
|  | **Strontium** |  |  |

1. **Complete the crossword below.**

|  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- |
| usestitle | | | | | | |
|  | | | | | **Across**  **5** He (6)  **6** Ca (7)  **8** O (6)  **9** Fe (4)  **10** Cu (6)  **12** K (9)  **15** Na (6)  **16** Au (4)  **18** H (8)  **19** Be (9)  **24** Cl (8)  **25** S (7)  **26** Ne (4) | |
| **Down**  **1** N (8)  **2** Ni (6)  **3** Al (9)  **4** As (7)  **7** Zn (4)  **11** P (10)   1. Pu (9)   **13** Ar (5) | | **14** Si (7)  **17** F (8)   1. B (5)   **20**  Pb (4)  **21** Hg (7)   1. Ag (6)   **23** C (6) | | atommm | | |
|  |  | |  | |  |  |